AQA

Please write clearly in block capitals.

Centre number | | | | | | Candidate number | | | | |

Surname _____

Forename(s) _____

Candidate signature _____

I declare this is my own work.

A-level CHEMISTRY

Paper 2 Organic and Physical Chemistry

Time allowed: 2 hours

Materials

For this paper you must have:
- the Periodic Table/Data Booklet, provided as an insert (enclosed)
- a ruler with millimetre measurements
- a scientific calculator, which you are expected to use where appropriate.

Instructions
- Use black ink or black ball-point pen.
- Fill in the boxes at the top of this page.
- Answer **all** questions.
- You must answer the questions in the spaces provided. Do **not** write outside the box around each page or on blank pages.
- If you need extra space for your answer(s), use the lined pages at the end of this book. Write the question number against your answer(s).
- All working must be shown.
- Do all rough work in this book. Cross through any work you do not want to be marked.

Information
- The marks for questions are shown in brackets.
- The maximum mark for this paper is 105.

For Examiner's Use	
Question	Mark
1	
2	
3	
4	
5	
6	
7	
8	
9	
10	
TOTAL	

Do not write outside the box

Answer **all** questions in the spaces provided.

0 1 Coconut oil contains a triester with three identical R groups.
This triester reacts with potassium hydroxide.

0 1 . 1 Complete the equation by drawing the structure of the other product of this reaction in the box.

Name the type of compound shown by the formula RCOOK

Give **one** use for this type of compound.

[3 marks]

Type of compound Carboxylate salt

Use Soap

0 1 . 2 The triester in coconut oil has a relative molecular mass, $M_r = 638.0$
In the equation shown at the start of Question **01**, R represents an alkyl group that can be written as $CH_3(CH_2)_n$

Deduce the value of n in $CH_3(CH_2)_n$
Show your working.

[3 marks]

$638 = 173 + 3(15 + 14n)$

Mr ester fragment = 173

$638 - (173 + 45) = 42n$

$n = \frac{418}{42} = 9.952880 \approx 10$

n 10

Do not write outside the box

0 1 . 3 A 1.450 g sample of coconut oil is heated with 0.421 g of KOH in aqueous ethanol until all of the triester is hydrolysed.
The mixture is cooled.
The remaining KOH is neutralised by exactly 15.65 cm^3 of 0.100 mol dm^{-3} HCl

Calculate the percentage by mass of the triester (M_r = 638.0) in the coconut oil.

[6 marks]

$n = \frac{m}{M_r}$ $\quad n_{HCl} = 0.100 \times 0.01565$

$n = cV$ $\quad = 1.565 \times 10^{-3}$ moles

$n_{KOH_{ini}} = \frac{0.421}{56.1} = 7.50 \times 10^{-3}$ moles

$n_{KOH_{used}} = 7.50 \times 10^{-3} - 1.565 \times 10^{-3} = 5.939 \times 10^{-3}$ moles

$n_{Ester} = \frac{5.939 \times 10^{-3}}{3} = 1.980 \times 10^{-3}$ moles

$mass_{Ester} = 1.980 \times 10^{-3} \times 638.0 = 1.263\,g$

$\%_{mass} = \frac{1.263}{1.450} \times 100 = 87.1\%$

Percentage by mass 87%

Do not write outside the box

0 1 . 4 Suggest why aqueous ethanol is a suitable solvent when heating the coconut oil with KOH.

Give a safety precaution used when heating the mixture.
Justify your choice.

[3 marks]

Reason Aqueous ethanol will dissolve both the oil and the KOH

Safety precaution Use a water bath to heat the reaction mixture

Justification Use of flames to heat the mixture risks igniting the ethanol.

15

Do not write outside the box

0 2 This question is about fuels.

0 2 . 1 The petrol fraction obtained from crude oil can be used as fuel in cars.

State the meaning of fraction, as used in the term petrol fraction.

[1 mark]

A group of compounds (in this case hydrocarbons) with similar boiling points

0 2 . 2 Hexadecane ($C_{16}H_{34}$) can be cracked at high temperature to form petrol.

Complete the equation to show the cracking of one molecule of hexadecane to form hexane and cyclopentane only.

Give the name of a catalyst used in this cracking reaction.

[3 marks]

$C_{16}H_{34} \rightarrow$ C_6H_{14} + $2C_5H_{10}$

Catalyst Zeolite

0 2 . 3 Carbon dioxide is formed when petrol is burned.
Carbon dioxide acts as a greenhouse gas when it absorbs infrared radiation.

Give a reason why carbon dioxide absorbs infrared radiation.

[1 mark]

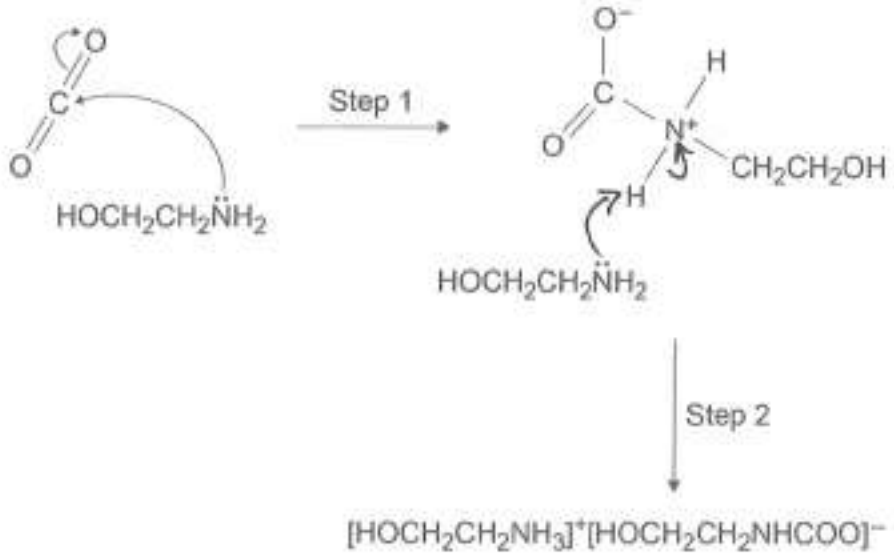
The C=O bond in CO_2 will vibrate at the same frequency as IR radiation.

Question 2 continues on the next page

0 2 . 4 Compound **Z** ($HOCH_2CH_2NH_2$) can be used to remove carbon dioxide from the mixture of waste gases produced in some power stations.

Figure 1 shows part of a suggested mechanism for the reaction of **Z** with carbon dioxide.

Figure 1



Draw **two** curly arrows to complete the mechanism in **Figure 1**.

Name compound **Z** ($HOCH_2CH_2NH_2$)

Deduce the role of **Z** in step **2** of the mechanism.

[4 marks]

Name 2-amino ethanol

Role Base

Do not write outside the box

0 2 . 5 $HOCH_2CH_2NH_2$ can be represented as XNH_2
$[HOCH_2CH_2NH_3]^+$ can be represented as $[XNH_3]^+$

Draw the shape of XNH_2 and of $[XNH_3]^+$

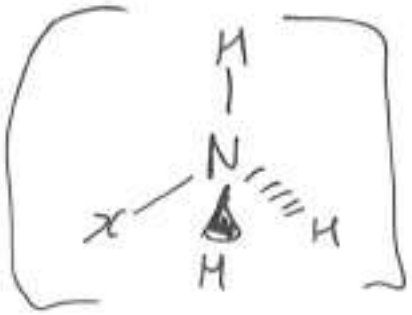
State whether the H–N–H bond angle in XNH_2 is greater than, the same as, or smaller than that in $[XNH_3]^+$

Explain your answer.

[4 marks]

Shape of XNH_2

Shape of $[XNH_3]^+$



Bond angle Smaller

Explanation lone pair repulsion is greater than bonding pair repulsion.

Question 2 continues on the next page

Do not write outside the box

0 2 . 6 Bioethanol is used as an alternative to fossil fuels.

This statement appeared on a website.

"The fact that bioethanol is a carbon-neutral fuel outweighs the environmental disadvantages of producing bioethanol."

Evaluate this statement.

In your answer you should include:
- an outline of how bioethanol is produced
- relevant equations
- analysis of the environmental impacts.

[6 marks]

Bioethanol is produced through the fermentation of plant based organic matter. This leads to the claim that bioethanol is carbon neutral. Photosynethis in the plants will take in CO_2 from the air during growth. These plants are then harvested, and the sugars produced during photosynthesis fermented to produce bio-ethanol with a side product of CO_2. This bioethanol can then be used in combustion as a fuel, producing futher CO_2. In specting the reactions at each stage does suggest that Bio-ethanol is a carbon neutral fuel.

1. Photosynthesis: Total CO_2 in = 6 CO_2

$$6CO_2 + 6H_2O \rightarrow C_6H_{12}O_6 + 6O_2$$

↑ sugar for fermentation.

2. Fermentation: CO_2 out = $2CO_2$

$$C_6H_{12}O_6 \xrightarrow{\text{Heat}} 2C_2H_5OH + 2CO_2$$

3. Combustion: CO_2 out = $4CO_2$

$$2C_2H_5OH + 6O_2 \longrightarrow 6H_2O + 4CO_2$$

Across fermentation we produce 2 moles of CO_2 and across combustion we produce 4. This gives a total of $6CO_2$ out. This is ballanced by the $6CO_2$ taken in during photo Synthesis. However, fermentation requires heat and so energy. This energy production will require fuel and as such produce more CO_2. Distribution of the bioethanol will also produce CO_2. As such, bioethanol is not Carbon neutral. Bio-ethanol also presents issues of deforestation for land to grow plants for fuel production as as a reduction in biodiversity. As such it can not be definitely stated that bioethanol is good because it is Carbon neutral and this outweight environmental concerns. Bioethanol is not Carbon neutral and so a more nuanced evaluation is required.

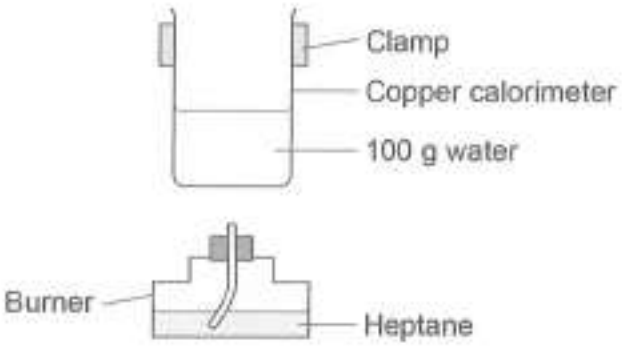
19

Do not write outside the box

0 3 A student does an experiment to determine a value for the enthalpy of combustion of heptane.

Figure 2 shows some of the apparatus used.

Figure 2



0 3 . 1 Design a table to record all the readings necessary to determine an experimental value for the enthalpy of combustion for heptane in this experiment.

[2 marks]

	Temp / °C		Mass / g
Initial		Burner Before	
Final		Burner After	
ΔT		Mass Heptane Burned	

0 3 . 2 The student considered using a glass beaker on a tripod and gauze instead of the clamped copper calorimeter.

Suggest **two** disadvantages of using a glass beaker on a tripod and gauze.

[2 marks]

Disadvantage 1 Glass is a poor conductor of heat

Disadvantage 2 The tripod and Gauze would reduce heat transfer by themselves absorbing some heat

Do not write outside the box

0 3 . 3 Suggest **two** reasons why the value of enthalpy of combustion from this experiment is less exothermic than a data book value.

[2 marks]

Reason 1 Heat loss to the surroundings

Reason 2 Combustion was not complete

0 3 . 4 Suggest **one** addition to this apparatus that would improve the accuracy of the enthalpy value obtained.

[1 mark]

Use a wind sheild

7

Turn over for the next question

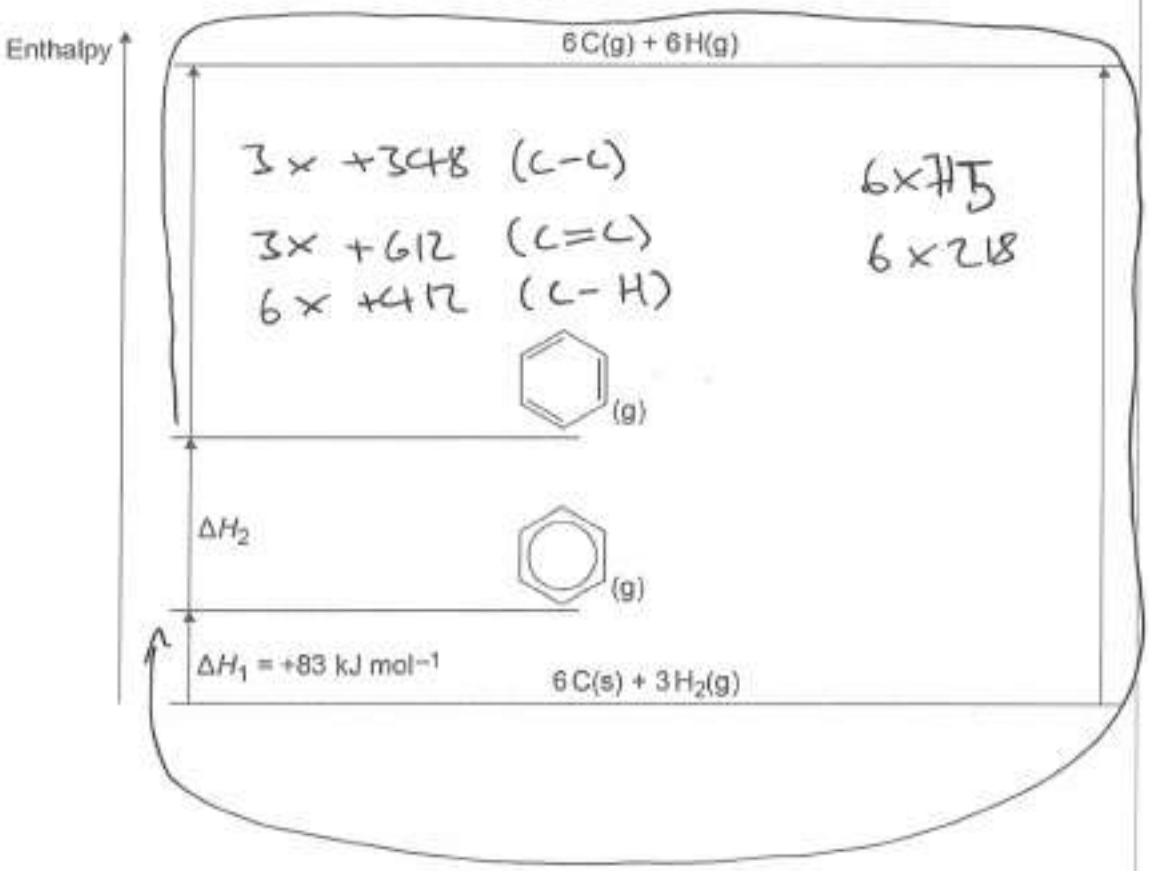
Do not write outside the box

0 4 Kekulé suggested this structure for benzene.

Benzene is now represented by this structure.

Figure 3 shows the relative stability of  compared to 

Figure 3



Do not write outside the box

0 4 . 1 Use **Figure 3** and the data shown in **Table 1** to calculate ΔH_2

[3 marks]

Table 1

	ΔH / kJ mol^{-1}
Enthalpy of atomisation for carbon	+715
Enthalpy of atomisation for hydrogen	+218
Bond enthalpy (C–C)	+348
Bond enthalpy (C=C)	+612
Bond enthalpy (C–H)	+412

Total atomisation energy = $(6 \times 715) + (6 \times 218) = 5598$ kJ mol^{-1}

Total bond breaking energy = $3(348) + 3(612) + 6(412) = 5352$ kJ mol^{-1}

$\Delta H_2 = 5352 -$

ΔH_2 _____ kJ mol^{-1}

0 4 . 2 Explain, in terms of structure and bonding, why 

is more thermodynamically stable than 

[1 mark]

The π electrons are delocalised around the ring

Turn over ▶

Do not write outside the box

0 4 . 1 Use **Figure 3** and the data shown in **Table 1** to calculate ΔH_2

[3 marks]

Table 1

	ΔH / kJ mol^{-1}
Enthalpy of atomisation for carbon	+715
Enthalpy of atomisation for hydrogen	+218
Bond enthalpy (C–C)	+348
Bond enthalpy (C=C)	+612
Bond enthalpy (C–H)	+412

Atomisation = $6(715) + 6(218) = 5598$ kJ mol^{-1}

Bond Breaking = $3(348) + 3(612) + 6(412) = 5352$ kJ mol^{-1}

$\Delta H_2 = 5598 - 5352 - 83$

$= +163$ kJ mol^{-1}

ΔH_2 + 163 kJ mol^{-1}

0 4 . 2 Explain, in terms of structure and bonding, why is more thermodynamically stable than

[1 mark]

π electrons are delocalised across the ring

Turn over ▶

Do not write outside the box

0 4 . 3 A mixture of concentrated nitric acid and concentrated sulfuric acid reacts with benzene.

Figure 4 shows the incomplete mechanism for this reaction.

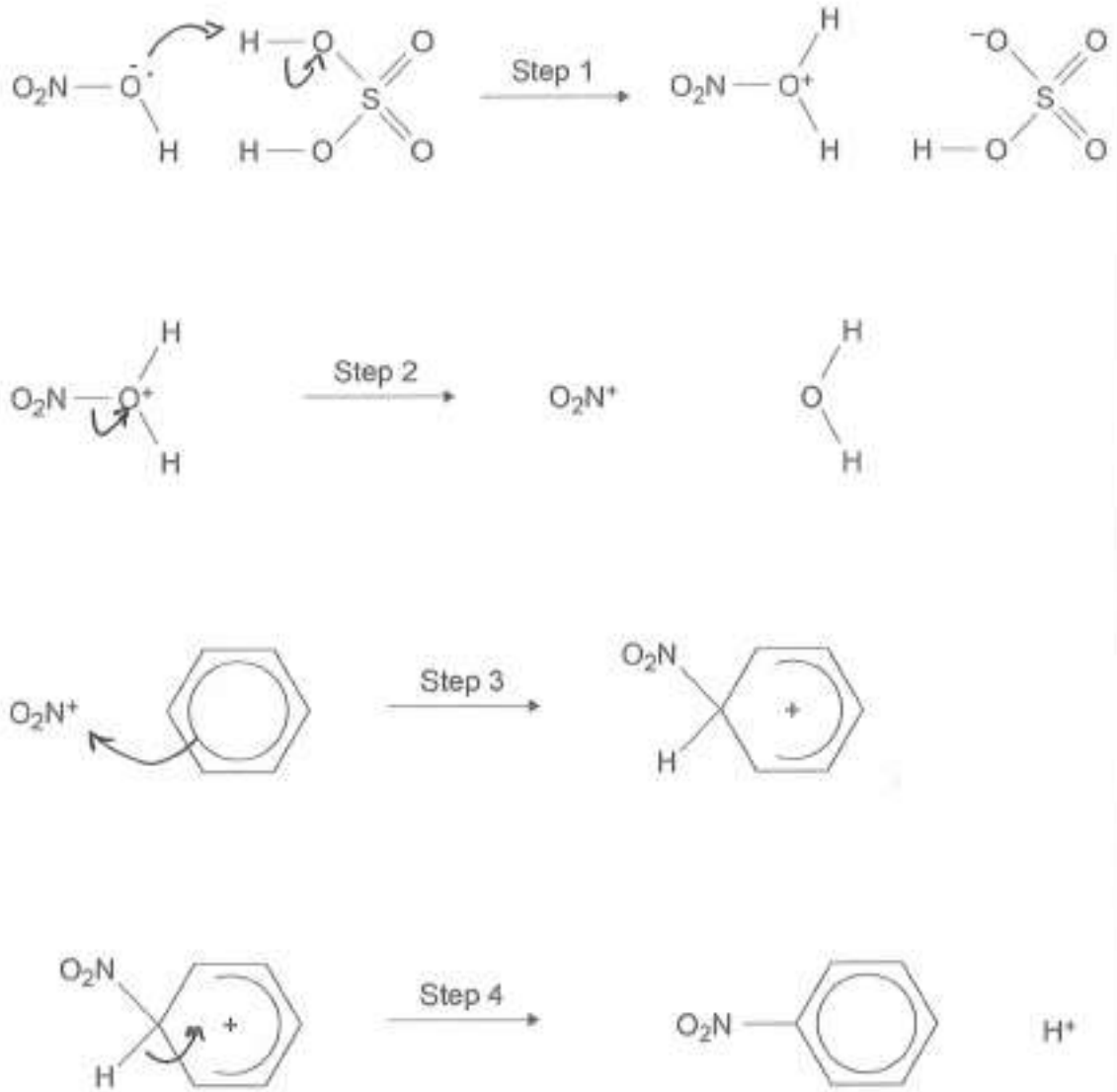
Name the mechanism.

Complete the mechanism in **Figure 4** by adding
- any lone pairs of electrons involved in each step
- **two** curly arrows in step **1**
- a curly arrow in step **2**
- a curly arrow in step **3**
- a curly arrow in step **4**.

[5 marks]

Name of mechanism Electrophilic Substitution.

Figure 4



9

Do not write outside the box

0 5 This question is about equilibrium.

0 5 . 1 1 mol of a diester with molecular formula $C_7H_{12}O_4$ is added to 1 mol of water in the presence of a small amount of catalyst.
The mixture is left to reach equilibrium at a constant temperature.

$$C_7H_{12}O_4(l) + 2\,H_2O(l) \rightleftharpoons 2\,CH_3COOH(l) + HO(CH_2)_3OH(l)$$

At equilibrium, x mol of ethanoic acid are present in the mixture.

Complete **Table 2** by deducing the amounts, in terms of x, of the diester, water and diol present in the equilibrium mixture.

[3 marks]

Table 2

	Amount in the mixture / mol			
	Diester	**Water**	**Acid**	**Diol**
At the start	1	1	0	0
At equilibrium	$1 - \frac{x}{2}$	$1 - x$	x	$\frac{x}{2}$

0 5 . 2 Deduce the structure of the diester in Question **05.1**

[1 mark]

Question 5 continues on the next page

Do not write outside the box

0 5 . 3 A new equilibrium mixture of the substances from Question **05.1** is prepared at a different temperature.

$$C_7H_{12}O_4(l) + 2H_2O(l) \rightleftharpoons 2CH_3COOH(l) + HO(CH_2)_3OH(l)$$

Table 3 shows the amount of each substance in this new equilibrium mixture.

Table 3

	Amount in the mixture / mol			
	Diester	**Water**	**Acid**	**Diol**
At equilibrium	0.971	To be calculated	0.452	0.273

The value of the equilibrium constant, K_c is 0.161 at this temperature.

Calculate the amount of water, in mol, in this new equilibrium mixture.
Show your working.

[3 marks]

$$K_c = \frac{0.452^2 \times 0.273}{0.971 \times n H_2O^2} = 0.161$$

$$n H_2O^2 = \frac{0.452^2 \times 0.273}{0.971 \times 0.161}$$

$$n H_2O = \sqrt{0.357}$$

$$= 0.597 \text{ moles}$$

Amount of water 0.597 mol

7

Do not write outside the box

0 6 This question is about isomers with the molecular formula $C_5H_{10}O$

0 6 . 1 Draw the skeletal formula of a branched chain aldehyde with molecular formula $C_5H_{10}O$ that is optically active.

[1 mark]

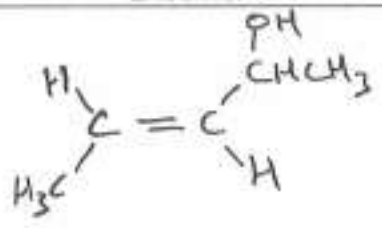
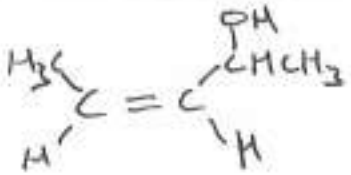
0 6 . 2 Describe how you distinguish between separate samples of the two enantiomers of the branched chain aldehyde $C_5H_{10}O$

[2 marks]

Use the rotation of plane polarised light. The two enantiomers will rotate plane polarised light in opposite directions

0 6 . 3 Draw the *E* and *Z* forms of a structural isomer of $C_5H_{10}O$ that shows **both** optical and geometric isomerism.

[2 marks]

E isomer	*Z* isomer
	

Question 6 continues on the next page

Do not write outside the box

0 6 . 4 Isomer J is cyclic and has an ether functional group (C–O–C)
Isomer J has only three peaks in its ^{13}C NMR spectrum.

Isomer J

Draw **two** other cyclic isomers of $C_5H_{10}O$ that have an ether functional group and only three peaks in their ^{13}C NMR spectra.

[2 marks]

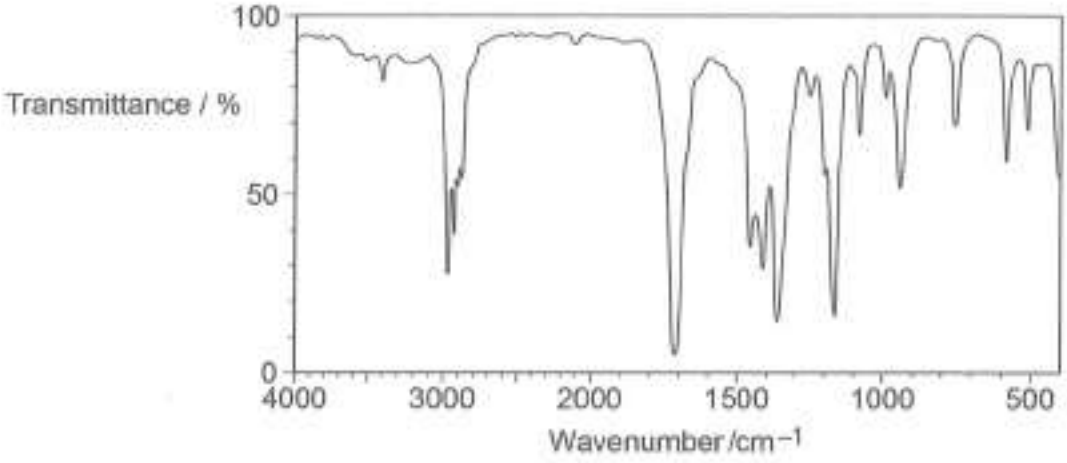
7

Do not write outside the box

0 7 This question is about spectroscopy.

0 7 . 1 Compound **K** has molecular formula C_4H_8O
Figure 5 shows the infrared spectrum of **K**.

Figure 5



Which functional group does **K** contain?

Tick (✓) **one** box.

[1 mark]

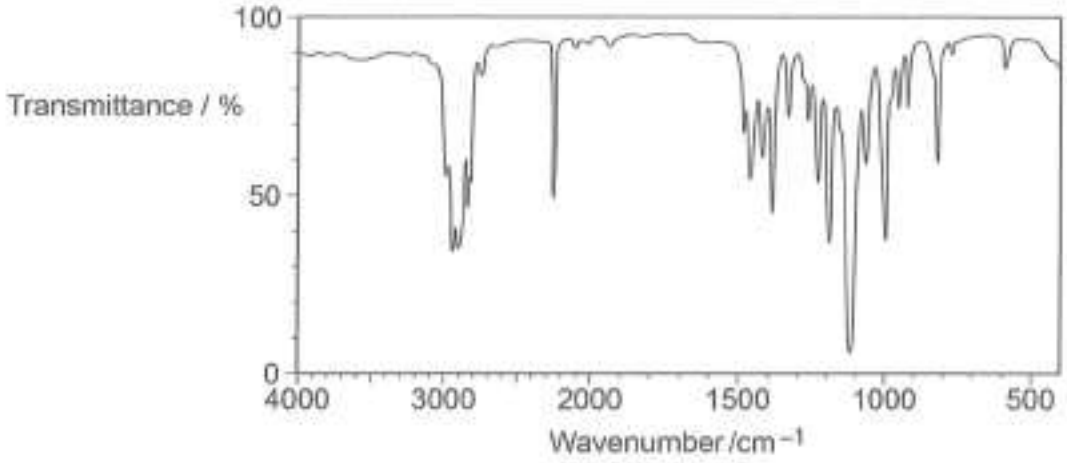
Functional Group				
alcohol	alkene	amine	carbonyl	nitrile
			✓	

Question 7 continues on the next page

Do not write outside the box

0 7 . 2 Compound **L** has molecular formula C_4H_7NO
Figure 6 shows the infrared spectrum of **L**.

Figure 6



L reacts with H_2 in the presence of a nickel catalyst to give compound **M**.

Suggest **three** ways in which the infrared spectrum of **M** is different from the infrared spectrum of **L**.

[3 marks]

1 The C≡N peak at 2220-2260 cm⁻¹ disappears

2 a peak at 3300-3500 cm⁻¹ appears for N-H

3 the fingerprint region will be different.

Do not write outside the box

0 7 . 3 **Figure 7** shows the ^{1}H NMR spectrum of **Q**, C_3H_7ClO

Figure 7

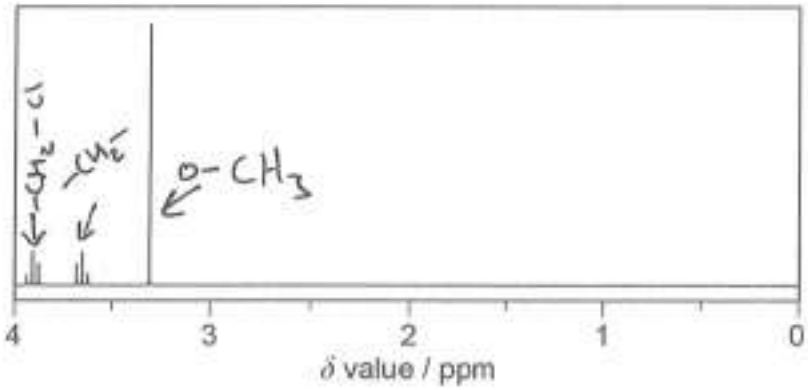


Table 4 shows the chemical shifts (δ values) and integration values for each peak.

Table 4

δ value / ppm	3.95	3.65	3.35
Integration value	0.6	0.6	0.9

Deduce the structure of **Q**.

Explain your answer.

[5 marks]

Structure:

$CH_3-O-CH_2CH_2Cl$

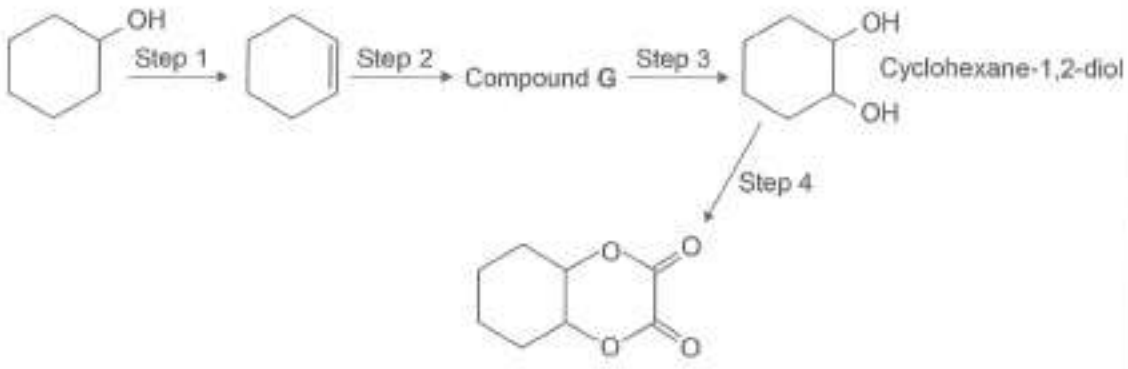
The integration ratio is 2:2:3 showing 2 sets of 2 hydrogens and one of 3. The peak at 3.95 is a triplet with an integration of 2, indicating CH_2-Cl next to a CH_2 group. The peak at 3.65 has an integration of 2 and indicates a $O-CH_2$ group next to a CH_2. The singlet at CH_3 3.35 indicates the CH_3 group with no adjacent H.

9

Turn over ▶

Do not write outside the box

0 8 This question is about making a diester from cyclohexanol.



0 8 . 1 State the type of reaction in step **1**.

Give the name of the reagent needed for step **1**.

[2 marks]

Type of reaction Dehydration

Reagent Concentrated H_2SO_4

0 8 . 2 State the reagents needed and give equations for step **2** and step **3**.

Show the structure of Compound **G** in your equations.

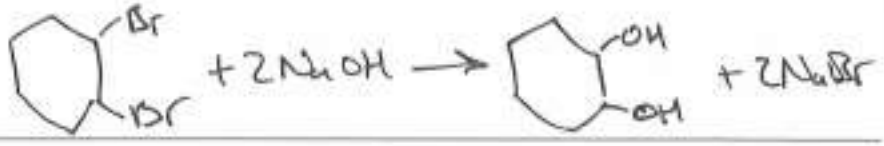
[4 marks]

Step 2 reagent Br_2

Step 2 equation

Step 3 reagent $NaOH$

Step 3 equation



Do not write outside the box

0 8 . 3 Cyclohexane-1,2-diol reacts with ethanedioyl dichloride.

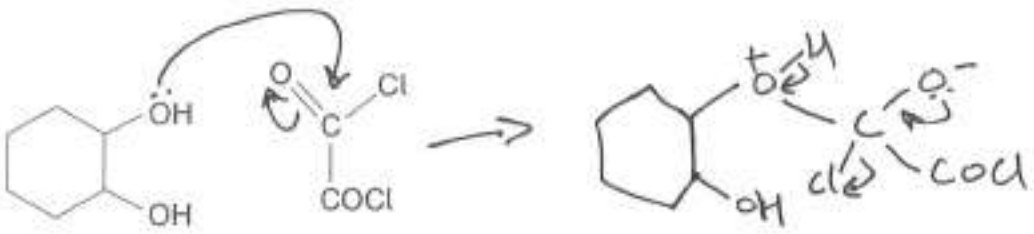
Give the name of the mechanism for this reaction.

Complete the mechanism to show the formation of **one** ester link in the first step of this reaction.

[5 marks]

Mechanism name Nucleophilic addition-elimination.

Mechanism



0 8 . 4 Suggest why chemists usually aim to design production methods

- with fewer steps
- with a high percentage atom economy.

[2 marks]

Fewer steps Improves yield by reducing loss between steps

High percentage atom economy less waste

13

Do not write outside the box

0 9 This question is about the ozone layer in the upper atmosphere.

0 9 . 1 State why the ozone layer is beneficial for living organisms.

[1 mark]

Absorbs harmful UV radiation

0 9 . 2 State how chlorofluorocarbons (CFCs) form chlorine atoms in the upper atmosphere.

[1 mark]

C – Cl bonds are broken

0 9 . 3 Give equations to show how chlorine atoms catalyse the decomposition of ozone.

[2 marks]

$Cl^{\bullet} + O_3 \rightarrow ClO^{\bullet} + O_2$

$ClO^{\bullet} + O_3 \rightarrow Cl^{\bullet} + 2O_2$

0 9 . 4 Hydrochlorofluorocarbons (HCFCs) have been used in place of CFCs.
In the mechanism to make an HCFC from a fluoroalkane, two incomplete steps are shown.

Complete each step in the mechanism.

Give the name of the type of step shown by both these equations.

[3 marks]

$Cl^{\bullet} + CH_2F_2 \rightarrow {}^{\bullet}CHF_2 + HCl$

${}^{\bullet}CHF_2 + Cl_2 \rightarrow CHClF_2 + Cl^{\bullet}$

Type of step propagation

7

Do not write outside the box

1 0 This question is about rates of reaction.
Iodine and propanone react together in an acid-catalysed reaction

$$CH_3COCH_3(aq) + I_2(aq) \rightarrow CH_3COCH_2I(aq) + HI(aq)$$

A student completed a series of experiments to determine the order of reaction with respect to iodine.

Method
- Transfer 25 cm^3 of 1.0 mol dm^{-3} propanone solution into a conical flask.
- Add 10 cm^3 of 1.0 mol dm^{-3} HCl(aq)
- Add 25 cm^3 of 5.0×10^{-3} mol dm^{-3} I_2(aq) and start a timer.
- At intervals of 1 minute, remove a 1.0 cm^3 sample of the mixture and add each sample to a separate beaker containing an excess of $NaHCO_3$(aq)
- Titrate the contents of each beaker with a standard solution of sodium thiosulfate and record the volume of sodium thiosulfate used.

1 0 . 1 Suggest why the 1.0 cm^3 portions of the reaction mixture are added to an excess of $NaHCO_3$ solution.

[2 marks]

The Sodium Hydrogencarbonate Solution will neutralise the acid Catalyst, Stopping the reaction

1 0 . 2 Suggest why the order of this reaction with respect to propanone can be ignored in this experiment.

[2 marks]

The Concentration of propanone is much larger than that of Iodine. As such the concentration of propanone is almost constant through out the reaction.

Question 10 continues on the next page

The volume of sodium thiosulfate solution used in each titration is proportional to the concentration of iodine in each beaker.

Table 5 shows the results of the experiment.

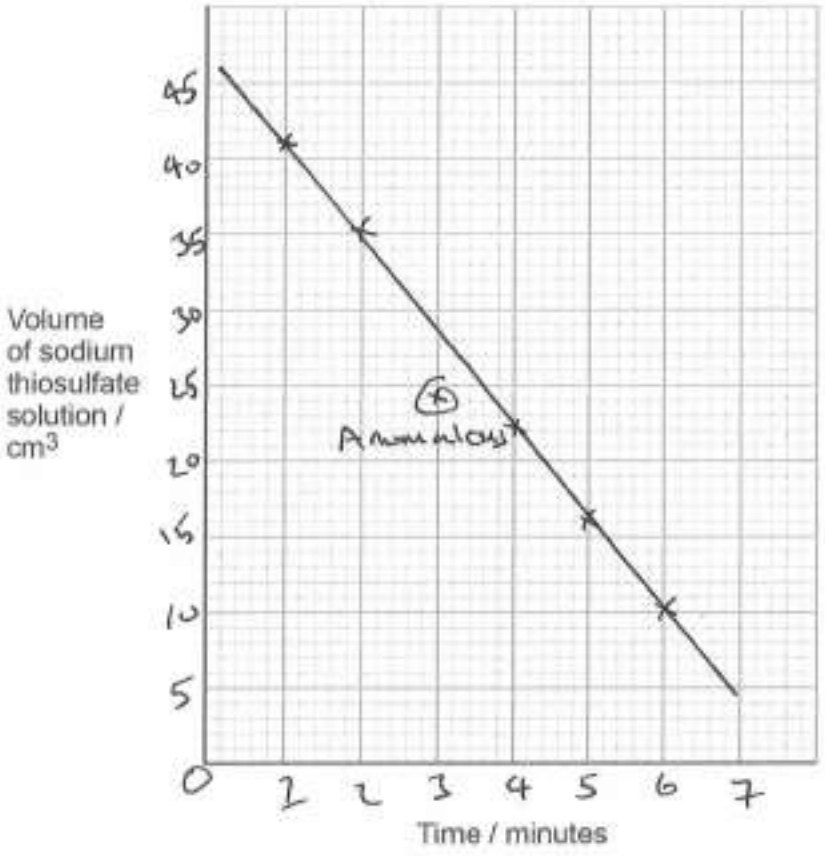
Table 5

Time / minutes	Volume of sodium thiosulfate solution / cm^3
1	41
2	35
3	24
4	22
5	16
6	10

1 0 . 3 Use the results in **Table 5** to draw a graph of volume of sodium thiosulfate solution against time.

Draw a line of best fit.

[3 marks]



Do not write outside the box

10.4 Explain how the graph shows that the reaction is zero-order with respect to iodine in the reaction between propanone and iodine.

[2 marks]

The graph produces a straight line plot. This means that the rate of reaction does not change with the concentration of iodine.

Question 10 continues on the next page

Do not write outside the box

Use **Figure 8** to calculate a value for the activation energy (E_a), in kJ mol^{-1}, for this reaction.

The gas constant $R = 8.31$ J K^{-1} mol^{-1}

[3 marks]

$m = \frac{\Delta y}{\Delta x}$ $\Delta y = (-14.1) - (-2.8)$

$= -11.3$

$\Delta x = 0.00180 - 0.00128$

$= 0.00052$

$m = \frac{-11.3}{0.00052} = -21731$ $m = -\frac{E_a}{R}$

$-E_a = mR$

$-E_a = -21731 \times 8.31$

$= 180,583 \text{ J mol}^{-1}$

$\ln k = \frac{-E_a}{R} \cdot \frac{1}{T} + \ln A$

$y = m \cdot x + c$

E_a 181 kJ mol^{-1}

12

END OF QUESTIONS