

**AQA, OCR, Edexcel**

**GCSE Science**

# **GCSE Chemistry**

Group 0 and 1

Answers

**M M E**

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Total Marks: /29

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Q1: What does the reaction of elements that form positive ions create? Circle one.

Metals

Non Metals

(1 mark)

Q2: Explain the differences between metals and non-metals when using electricity.

A= 1 mark for each of the following:

- Metals – Conduct
- Non metals – Insulators / non conductors

(2 marks)

Q3: What happens when elements from group 7 gain electrons?

A= Form negative ions

(1 mark)

Q4: Group 0 contains the noble gasses. Give 2 examples of a noble gas.

A= Accept any 2 of the following:

- Helium
- Neon
- Argon
- Krypton
- Xenon
- Radon

(2 marks)

Q5: Define the properties of a noble gas.

A= Accept any 2 of the following:

- Unreactive
- Stable arrangement of electrons
- Don't easily form molecules

(2 marks)

Q6: Predict what will happen to the boiling points of the Nobel gasses going down the periodic group.

A = Boiling point increases

(1 mark)

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Q7: Alkaline metals are part of which group? Circle one.

Group 1

Group 5

Group 6

(1 mark)

Q8: How does the electron structure of the group 1 elements give their characteristic properties?

A= Single electron in their outer shell

(1 mark)

Q9: How are group 1 metals stored and why?

A= 1 mark for each of the following points:

- Stored in oil
- Very reactive

(2 marks)

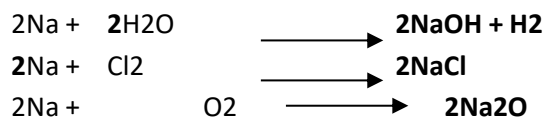
Q10: Discuss the properties of the group 1 metals

A= Accept any 6 of the following:

- Low Density
- Float on water
- Very soft
- Very reactive
- Reactive with oxygen
- Burn vigorously
- Low boiling point
- Low melting point
- Reactive with water

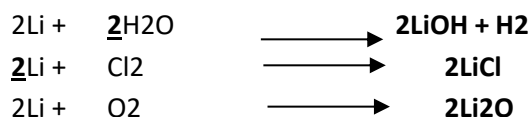
(6 marks)

Q11: Complete the following Equations to show sodium's reactivity with water, chlorine and oxygen.



(3 marks)

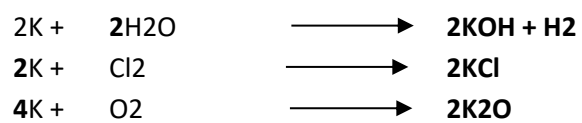
Q12: Complete the following equations to show lithium's reactivity with water chlorine and oxygen.



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(3 marks)

Q13: Complete the following equations to show potassium's reactivity with water, chlorine and oxygen.



(3 marks)

Q14: How does the reactivity of the alkali metals change going down the periodic group?

A= Increases

(1 mark)