

Please check the examination details below before entering your candidate information

Candidate surname

Other names

Centre Number

Candidate Number

**Pearson Edexcel Level 1/Level 2 GCSE (9–1)**

**Friday 17 May 2024**

Morning (Time: 1 hour 10 minutes)

Paper  
reference

**1SC0/1CH**

**Combined Science**

**PAPER 2**

**Higher Tier**

**You must have:**

Calculator, ruler, Periodic Table (enclosed)

Total Marks

### Instructions

- Use **black** ink or ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided – *there may be more space than you need.*
- Calculators may be used.
- Any diagrams may NOT be accurately drawn, unless otherwise indicated.
- You must **show all your working out** with **your answer clearly identified** at the **end of your solution**.

### Information

- The total mark for this paper is 60.
- The marks for **each** question are shown in brackets – *use this as a guide as to how much time to spend on each question.*
- In questions marked with an **asterisk** (\*), marks will be awarded for your ability to structure your answer logically, showing how the points that you make are related or follow on from each other where appropriate.

### Advice

- Read each question carefully before you start to answer it.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ►

P74438A

©2024 Pearson Education Ltd.  
F:1/1/1/1/



  
Pearson

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

**BLANK PAGE**



DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

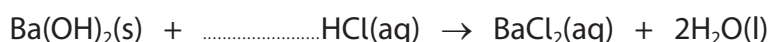
Answer ALL questions. Write your answers in the spaces provided.

Some questions must be answered with a cross in a box ☒. If you change your mind about an answer, put a line through the box ☒ and then mark your new answer with a cross ☒.

1 Barium hydroxide reacts with dilute hydrochloric acid to form barium chloride solution and water.

(a) (i) Complete the balanced equation for the reaction by adding a **number** in front of HCl(aq).

(1)



(ii) State what you would **see** during the reaction.

(1)

.....  
.....

(b) A student investigated how the pH of the mixture changed as barium hydroxide was added to dilute hydrochloric acid.

The student used this method.

**step 1** measure out 50 cm<sup>3</sup> of dilute hydrochloric acid into a beaker using a measuring cylinder

**step 2** use a glass rod to place a drop of the acid onto a piece of universal indicator paper and record the pH

**step 3** add one spatula measure of barium hydroxide to the acid in the beaker and stir

**step 4** use the glass rod to place a drop of the mixture onto a new piece of universal indicator paper and record the pH again

**step 5** repeat steps 3 and 4 until there is no further change in the pH.



(i) Name a piece of equipment that could be used to measure the pH of a substance more accurately than universal indicator paper.

(1)

(ii) Explain why, in step 3, the mixture was stirred after adding the barium hydroxide.

(2)

(iii) Figure 1 shows the student's results.

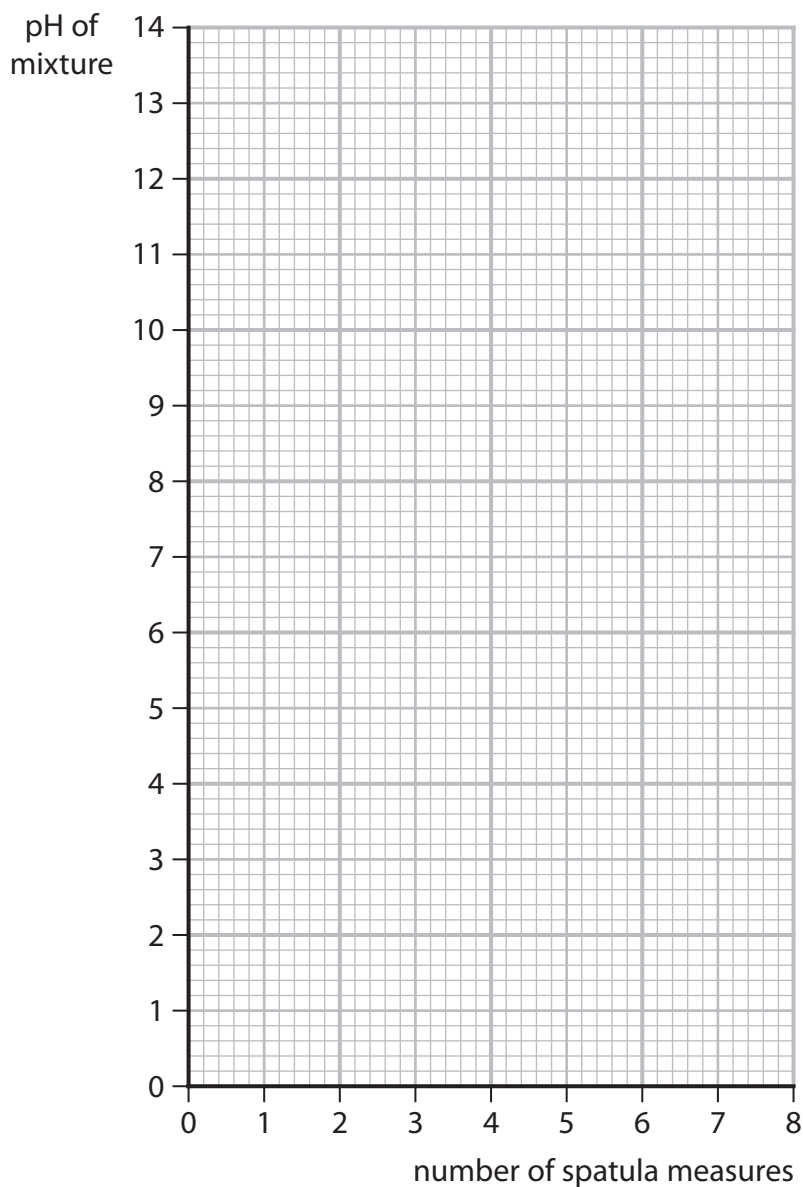
number of spatula measures of barium hydroxide	pH of mixture
0	1
1	1
2	1
3	1
4	3
5	8
6	12
7	13
8	13

Figure 1



Plot a graph of the pH of the mixture against the number of spatula measures of barium hydroxide.

(3)



(iv) Use the graph to find the pH of the mixture when 4.5 spatula measures of barium hydroxide are added.

(1)

pH of the mixture = .....

**(Total for Question 1 = 9 marks)**

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA



DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

**BLANK PAGE**



DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

2 Sodium carbonate has the formula  $\text{Na}_2\text{CO}_3$ .

(a) Sodium carbonate contains  $\text{Na}^+$  ions and  $\text{CO}_3^{2-}$  ions.

(i) The atomic number of sodium is 11.

What is the electronic configuration of the  $\text{Na}^+$  ion?

(1)

- A 1
- B 2.8
- C 2.8.1
- D 2.8.2

(ii) Explain why solid sodium carbonate **cannot** conduct electricity but a solution of sodium carbonate **can** conduct electricity.

(3)

.....

.....

.....

.....

.....

.....

.....

(b) Calculate the percentage by mass of sodium in sodium carbonate,  $\text{Na}_2\text{CO}_3$ .

$$\text{percentage by mass of element} = \frac{\text{total relative atomic mass of element}}{\text{relative formula mass of compound}} \times 100$$

(relative atomic masses: C = 12, O = 16, Na = 23)

(3)

.....

.....

.....

.....

percentage by mass of sodium = .....

**(Total for Question 2 = 7 marks)**



3 This question is about the extraction of metals.

(a) Give **two** advantages of obtaining metals by recycling rather than by extracting them from their metal ores.

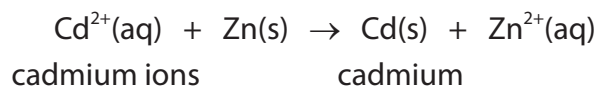
(2)

1 .....

2 .....

(b) (i) Small amounts of some metals are extracted using displacement reactions.

In one process, zinc dust is used to precipitate cadmium metal from a solution containing cadmium ions.



Explain why this displacement reaction can be described as a **redox reaction**.

(3)

.....  
.....  
.....  
.....  
.....  
.....  
.....  
.....  
.....  
.....



- (ii) The formula of the cadmium ion is  $\text{Cd}^{2+}$ .  
The formula of the phosphate ion is  $\text{PO}_4^{3-}$ .

Which is the formula of cadmium phosphate?

(1)

- A  $\text{Cd}_2(\text{PO}_4)_3$
- B  $\text{Cd}_3\text{PO}_{12}$
- C  $\text{Cd}_3(\text{PO}_4)_2$
- D  $\text{Cd}_3\text{P}_2\text{O}_8$

- (c) One of the alternative biological methods of extracting metals from very low-grade ores is bioleaching using bacteria.

Give one **disadvantage** of this method of extracting metals from low-grade ores.

(1)

.....

.....

- (d) Lead is low in the reactivity series.

Describe how to obtain a sample of lead from some lead oxide in the laboratory.

(2)

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

**(Total for Question 3 = 9 marks)**

DO NOT WRITE IN THIS AREA

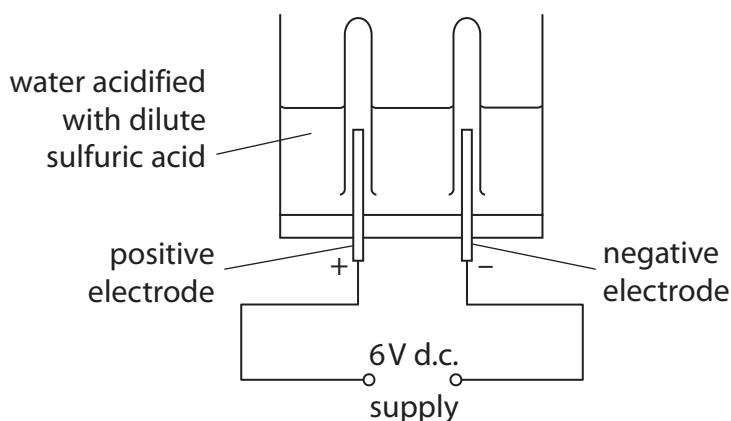
DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA



- 4 (a) Water, acidified with dilute sulfuric acid, was electrolysed for 10 minutes using inert electrodes.

Figure 2 shows the apparatus used.



**Figure 2**

- (i) In this electrolysis, the acidified water is an electrolyte.

Explain why acidified water is an electrolyte.

(2)

.....

.....

.....

.....

- (ii) Hydrogen collects at the negative electrode and oxygen collects at the positive electrode.

Which of these shows the results after 10 minutes of this electrolysis?

(1)

<input type="checkbox"/> <b>A</b>		<input type="checkbox"/> <b>B</b>	
<input type="checkbox"/> <b>C</b>		<input type="checkbox"/> <b>D</b>	



(iii) Complete and balance the half equation for the formation of oxygen at the positive electrode in this electrolysis. (2)



(b) Copper sulfate solution was electrolysed for 10 minutes using copper electrodes.

Figure 3 shows the mass of the cathode and the appearance of the copper sulfate solution before electrolysis and after electrolysis.

	mass of cathode in g	appearance of copper sulfate solution
before electrolysis	5.32	pale blue solution
after electrolysis	5.87	pale blue solution

Figure 3

(i) Describe what should be done to the copper cathode, after it has been removed from the copper sulfate solution, before its final mass is determined. (2)

.....

.....

.....

.....

(ii) Explain, in terms of ions, the change in mass of the cathode shown in Figure 3. (2)

.....

.....

.....

.....

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA



P 7 4 4 3 8 A 0 1 1 2 0

(iii) Explain why the appearance of the copper sulfate solution did not change during the electrolysis.

(2)

.....

.....

.....

.....

**(Total for Question 4 = 11 marks)**

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA



DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

**BLANK PAGE**



P 7 4 4 3 8 A 0 1 3 2 0

5 This question is about the properties of different substances.

(a) Silicon tetrachloride is a simple molecular covalent compound.

(i) A molecule of silicon tetrachloride is composed of a silicon atom and four chlorine atoms.

- a silicon atom has 4 outer electrons
- a chlorine atom has 7 outer electrons

Draw a dot and cross diagram of a molecule of silicon tetrachloride,  $\text{SiCl}_4$ .

Show outer electrons only.

(2)

(ii) Explain why simple molecular covalent compounds such as silicon tetrachloride have low melting and boiling points.

(2)

.....

.....

.....

.....



(b) Part of the structure of rubidium bromide is shown in Figure 4.

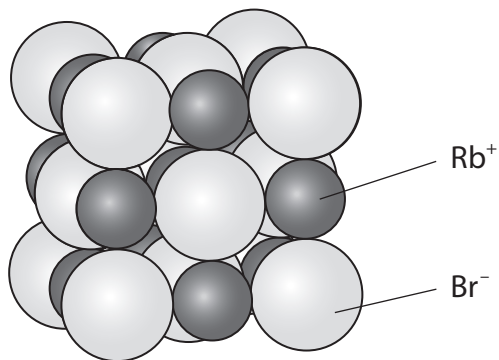


Figure 4

Which row shows the most likely melting point and boiling point of rubidium bromide?

(1)

	melting point in °C	boiling point in °C
<input type="checkbox"/> A	6.93	134.0
<input type="checkbox"/> B	69.3	134.0
<input type="checkbox"/> C	69.3	1340
<input type="checkbox"/> D	693	1340

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA



\*(c) Diamond and graphite are two forms of carbon.

Figure 5 shows how the carbon atoms are arranged in a part of the structure of each of these forms of carbon.

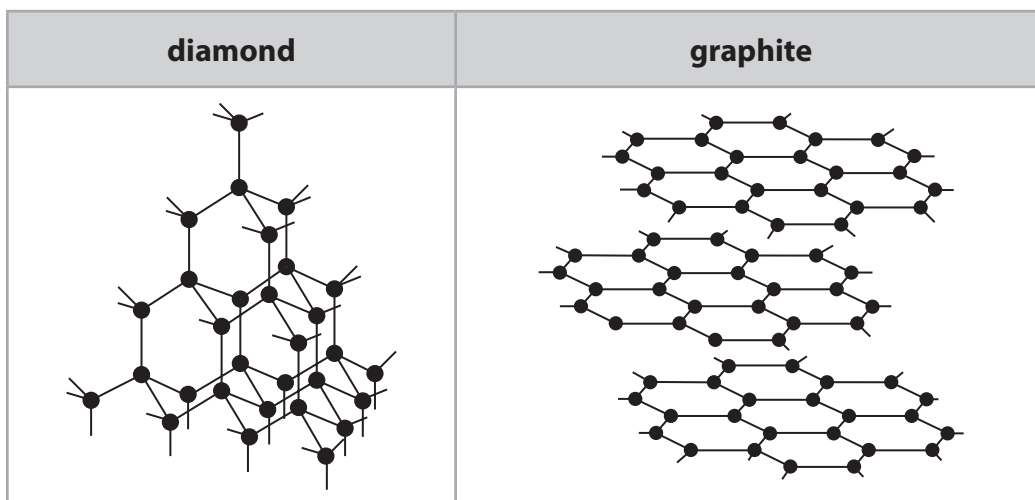


Figure 5

- diamond is one of the hardest known substances on Earth and is used in cutting tools.
- graphite is soft and flaky.
- diamond is a poor electrical conductor, but graphite is a good electrical conductor.

Explain, in terms of structure and bonding, these properties of diamond and graphite.

(6)

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....



DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

Handwriting practice area with 30 horizontal dotted lines.

**(Total for Question 5 = 11 marks)**



- 6 (a) An investigation was carried out on the reactivity of four metals, **D**, **E**, **F** and **G**.

Equal sized pieces of these metals were placed in excess dilute hydrochloric acid and left for three minutes.

Figure 6 shows the observations of the reactions for metals **D**, **E** and **F**.

metal	observations with dilute hydrochloric acid
<b>D</b>	Bubbles formed quickly. After three minutes all the metal had reacted.
<b>E</b>	Bubbles formed very quickly. No metal remaining after three minutes.
<b>F</b>	A few bubbles were seen to form. The metal looked unchanged after three minutes.
<b>G</b>	

Figure 6

The order of reactivity for these metals is shown in Figure 7.

<b>E</b>	<b>D</b>	<b>G</b>	<b>F</b>
----------	----------	----------	----------

most reactive least reactive

Figure 7

- (i) Use the information in Figure 6 and Figure 7 to suggest the observations that would be made for metal **G**.

(2)

.....

.....

.....

.....



(ii) The dilute hydrochloric acid used in this reaction is a strong acid.

Explain the meaning of the terms **dilute** and **strong acid**.

(4)

dilute

.....

.....

.....

.....

strong acid

.....

.....

.....

.....

(b) The formula of lead ethanoate is  $\text{Pb}(\text{CH}_3\text{COO})_2$ .

Calculate the number of **atoms** that combine together to form 16.25 g of lead ethanoate.

(relative atomic masses: H = 1.00, C = 12.0, O = 16.0, Pb = 207

Avogadro number =  $6.02 \times 10^{23}$ )

(4)

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

number of atoms = .....

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA

DO NOT WRITE IN THIS AREA



P 7 4 4 3 8 A 0 1 9 2 0

(c) Iron is more reactive than copper.

Iron will displace copper from copper nitrate solution.

Two possible balanced equations for the reaction are



It was found that 2.24 g of iron reacted with excess copper nitrate solution to form 3.81 g of copper.

Carry out a calculation, using the information above, to show which equation represents the reaction taking place.

(relative atomic masses: Fe = 56.0, Cu = 63.5)

(3)

.....

.....

.....

.....

.....

.....

.....

**(Total for Question 6 = 13 marks)**

**TOTAL FOR PAPER = 60 MARKS**



**Pearson Edexcel Level 1/Level 2 GCSE (9–1)**

**Friday 17 May 2024**

Paper  
reference

**1SC0/1CH**

**Combined Science**

**PAPER 2**

**Higher Tier**

**Periodic Table Insert**

**Do not return this Insert with the question paper**

*Turn over* ►

**P74438A**

©2024 Pearson Education Ltd.  
F:1/1/1/1/1/



P 7 4 4 3 8 A



**Pearson**

# The periodic table of the elements

1	2	3	4	5	6	7	0										
7 <b>Li</b> lithium 3	9 <b>Be</b> beryllium 4	11 <b>Na</b> sodium 11	12 <b>Mg</b> magnesium 12	13 <b>Al</b> aluminium 13	14 <b>N</b> nitrogen 7	15 <b>P</b> phosphorus 15	16 <b>O</b> oxygen 8	17 <b>F</b> fluorine 9	18 <b>Ne</b> neon 10								
19 <b>K</b> potassium 19	20 <b>Ca</b> calcium 20	21 <b>Sc</b> scandium 21	22 <b>Ti</b> titanium 22	23 <b>V</b> vanadium 23	24 <b>Cr</b> chromium 24	25 <b>Mn</b> manganese 25	26 <b>Fe</b> iron 26	27 <b>Co</b> cobalt 27	28 <b>Ni</b> nickel 28	29 <b>Cu</b> copper 29	30 <b>Zn</b> zinc 30	31 <b>Ga</b> gallium 31	32 <b>Ge</b> germanium 32	33 <b>As</b> arsenic 33	34 <b>Se</b> selenium 34	35 <b>Br</b> bromine 35	36 <b>Kr</b> krypton 36
37 <b>Rb</b> rubidium 37	38 <b>Sr</b> strontium 38	39 <b>Y</b> yttrium 39	40 <b>Zr</b> zirconium 40	41 <b>Nb</b> niobium 41	42 <b>Mo</b> molybdenum 42	43 <b>Tc</b> technetium 43	44 <b>Ru</b> ruthenium 44	45 <b>Rh</b> rhodium 45	46 <b>Pd</b> palladium 46	47 <b>Ag</b> silver 47	48 <b>Cd</b> cadmium 48	49 <b>In</b> indium 49	50 <b>Sn</b> tin 50	51 <b>Sb</b> antimony 51	52 <b>Te</b> tellurium 52	53 <b>I</b> iodine 53	54 <b>Xe</b> xenon 54
55 <b>Cs</b> caesium 55	56 <b>Ba</b> barium 56	57 <b>La*</b> lanthanum 57	72 <b>Hf</b> hafnium 72	73 <b>Ta</b> tantalum 73	74 <b>W</b> tungsten 74	75 <b>Re</b> rhenium 75	76 <b>Os</b> osmium 76	77 <b>Ir</b> iridium 77	78 <b>Pt</b> platinum 78	79 <b>Au</b> gold 79	80 <b>Hg</b> mercury 80	81 <b>Tl</b> thallium 81	82 <b>Pb</b> lead 82	83 <b>Bi</b> bismuth 83	84 <b>Po</b> polonium 84	85 <b>At</b> astatine 85	86 <b>Rn</b> radon 86

1  
**H**  
hydrogen  
1

## Key

relative atomic mass  
atomic symbol  
name  
atomic (proton) number

\* The elements with atomic numbers from 58 to 71 are omitted from this part of the periodic table.

*The relative atomic masses of copper and chlorine have not been rounded to the nearest whole number.*

