

AQA, OCR, Edexcel

GCSE Science

GCSE Chemistry

Group 0 and 1
Questions

M M E

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Total Marks: /29

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Q1: What does the reaction of elements that form positive ions create? Circle one.

Metals

Non Metals

(1 mark)

Q2: Explain the differences between metals and non-metals when using electricity.

(2 marks)

Q3: What happens when elements from group 7 gain electrons?

(1 mark)

Q4: Group 0 contains the noble gasses. Give 2 examples of a noble gas.

(2 marks)

Q5: Define the properties of a noble gas.

(2 marks)

Q6: Predict what will happen to the boiling points of the Nobel gasses going down the periodic group.

(1 mark)

Q7: Alkaline metals are part of which group? Circle one.

Group 1

Group 5

Group 6

(1 mark)

Q8: How does the electron structure of the group 1 elements give their characteristic properties?

(1 mark)

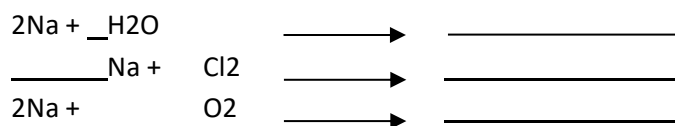
Q9: How are group 1 metals stored and why?

(2 marks)

Q10: Discuss the properties of the group 1 metals

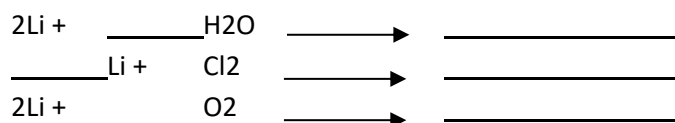
(6 marks)

Q11: Complete the following Equations to show sodium's reactivity with water, chlorine and oxygen.



(3 marks)

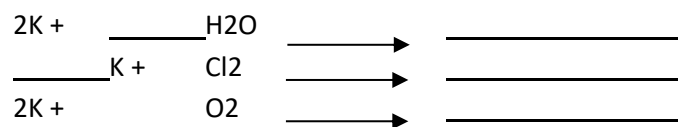
Q12: Complete the following equations to show lithium's reactivity with water chlorine and oxygen.



(3 marks)

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Q13: Complete the following equations to show potassium's reactivity with water, chlorine and oxygen.



(3 marks)

Q14: How does the reactivity of the alkali metals change going down the periodic group?

(1 mark)